**Interpreting equilibrium constant expressions (Level 2) examiners tips: Read these please!**

• K = products (being above the line, known as the numerator)

reactants (being below the line, known as the denominator)

• if K is large (> 1) then there will be more products than reactants so right hand/products side is favoured

• if K is small (< 1) then there will be more reactants than products so the left hand/reactants side is favoured

• the ONLY factor affecting the value of K is temperature  
• if the temperature INcreases the reaction proceeds towards the ENdothermic side (think ENTRANCE)...  
 because the endothermic side absorbs the heat energy

**Also…don’t be daft!**K only provides information on how far a reaction proceeds  
K doesn't provide any information on the rate of the reaction

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